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Chapter 4 Atomic

Structure33 SECTION

4.1 DEFINING THE

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ATOM (pages 101–103)

This section describes early atomic theories of matter and provides ways to understand the tiny size of individual atoms. Early Models of the Atom (pages 101–102) 1.

Democritus, who lived in Greece during the fourth century B.C.,

Name Date Class

ATOMIC STRUCTURE

4

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Structure Summary 4.1

Studying Atoms • The
ancient Greek

philosopher

Democritus believed
that all matter

consisted of extremely
small particles that
could not be divided. •

Aristotle did not think
there was a limit to the
number of times
matter could be ...

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Structure Summary 4.1

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Structure Section 4.1

Studying Atoms (pages

100-105) This section

discusses the

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development of atomic models.

Answers

Chapter 4: Atomic Structure Section

4.1: Studying Atoms

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Structure. WordWise.

Solve the clues to

determine which

vocabulary terms from

Chapter 4 are hidden in

the puzzle. Then find

and circle the terms in

the puzzle. The terms

may occur vertically,

horizontally, or

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diagonally. em a s s n
umb e run nl o r b i t a l
x a p i r e e n l t p t s p b
k s g n a c a s r d c r h l
o b e l d t g o f l s g a t l r
t s o r t g r n b t o n g z
b m o o p l q d c p p y q p
i u n n m a s s e s l n m c n
n u c l e u s t e u e n d r i o l
k m r r v c l u s v a b t o p k
z e l x m t w e ...

Chapter 4 Atomic Structure WordWise

Chapter 4 Atomic

Structure Section 4.3

Modern Atomic Theory

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(pages 113-118) This section focuses on the arrangement and behavior of electrons in atoms.

4.3 Modern Atomic Theory.pdf - Google Docs

Chapter 4: Atomic Structure Vocabulary Review. the sum of the protons and neutrons in the nucleus of that atom. atoms of the same element that have different numbers

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Chapter 4 Atomic Structure

of neutrons and different mass numbers. This activity was created by a Quia Web subscriber.

Quia - Chapter 4: Atomic Structure Vocabulary Review

Chapter 4- Atomic Structure Basics:
Notes, Review Quiz (Prentice Hall)
Tutorials: Simulations:
Simulations (U Colorado): Wave on a String, Making Waves,

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Radio Waves/EM, Wave
Interference,

Rutherford Experiment,
Hydrogen Atom Models

**Chemistry I - Mr.
Benjamin's
Classroom**

Start studying Section
4.1 Studying Atoms.

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Atoms Flashcards |**

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Quizlet

104 Chapter 4 • The
Structure of the Atom

■ Figure 4.2 In his book
A New System of
Chemical

Philosophy, John Dalton
presented his symbols
for the elements known
at that time and their
possible combinations.

Chapter 4: The Structure of the Atom

the number of orbitals
in energy levels 1

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Structure

through 4 What

mathematical

expression can you use

to calculate the

unknown? Level 1:

Level 3: Level 2: Level

4: 3. Look Back and

Check Is your answer

reasonable? The ratio

is the same for all four

energy levels. Also,

each orbital can

contain only two

electrons. Math

Practice

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**Structure Electrons
and Orbitals**

Chapter 4 Atomic

Structure WordWise 10

Terms. Emily__Guffey.

Chapter 4 Vocab (12

words) 12 Terms.

seversonm. chapter 4

vocab 12 Terms.

max220. chapter 4 40

Terms. panzerj140.

OTHER SETS BY THIS

CREATOR. chapter 12

neural tissue STUDY

GUIDE 50 Terms.

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Terms.

Workbook

Science Wordwise

chapter 4 Flashcards

| Quizlet

Section 4.2 The Structure of an Atom (pages 108–112) This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Chapter 4 Atomic Structure Section

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4.1 Studying Atoms

Atomic Structure - Unit

Test Paper 4. Q.1.

Select the correct answer from the answers in brackets to complete each sentence. Question 1.

An element has electronic configuration 2, 8, 1 and 12 neutrons. Its mass no. is ____ [11 / 23/ 12]

Answer: An element has electronic configuration 2, 8, 1 and 12 neutrons. Its

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mass no. is 23.

Question 2.

Answers
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Chemistry Class 9
ICSE Solutions
Atomic ...**

The atom of an element is made up of 4 protons, 5 neutrons and 4 electrons. What is its atomic number and mass number?

Answer: Protons = 4,
neutrons = 5, electrons
= 4 Atomic number =
4, Mass number = 4 +

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5 = 9. Question 16. (a) What are the two main parts of which an atom is made of? (b) Where is the nucleus of an atom situated ?

Selina Concise Chemistry Class 8 ICSE Solutions - Atomic ...

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Measuring Atomic Mass
Instead of grams, the
unit we use is the

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Atomic Mass Unit

(amu) It is defined as one-twelfth the mass of a carbon-12 atom.

Carbon-12 chosen because of its isotope purity. Each isotope has its own atomic mass, thus we determine the average from percent abundance.

Pearson Education
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the Atom - ppt

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Study Workbook

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